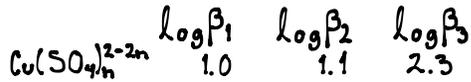


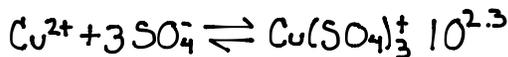
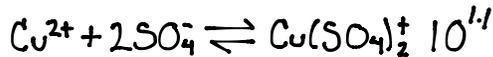
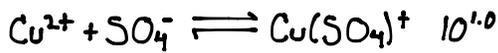
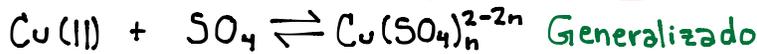
Si se mezclan 20 mL de una solución de Cu^{2+} 0.3 M con 40 mL de una solución de SO_4^{2-} 0.7 M. Escriba las reacciones que ocurren (ayúdate de una escala de predicción de reacciones).



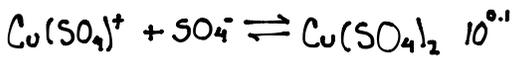
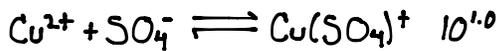
i) $(20\text{ mL})(0.3\text{ M}) = 6\text{ mmol}$ $(40\text{ mL})(0.7\text{ M}) = 28\text{ mmol}$

forma) $C_e = \frac{22\text{ mmol}}{60\text{ mL}} = 0.366\text{ M}$ 6 mmol

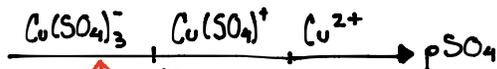
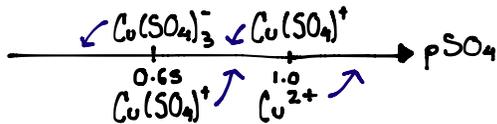
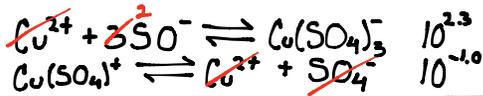
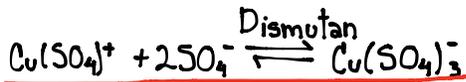
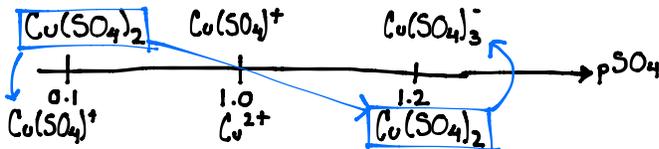
$$-\log(0.366\text{ M}) = 0.44$$



} Eq. Globales



} Eq. Sucesivos



0.44

.65

1.0

